**République Algérienne Démocratique et Populaire**

**Ministère de l’Enseignement Supérieur et de la Recherche Scientifique**

**Université Abderrahmane Mira de Béjaia**

**Faculté de Technologie**

**Département de Technologie 1ère année**

**MODULE : TP de Chimie 1**

***Compte rendu***

**TP**

**N° 3**

**Dosage acido-basique**

**Noms des étudiants : Groupe :**

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| ***Note*** | ***……… / 20*** |

**Date :**

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* Donner sous forme d’un tableau, les couleurs de chaque solution après l’ajout du bleu de bromothymol et de phénolphtaléine. Préciser ensuite le caractère acido-basique ainsi que leurs réactions en milieu aqueux.

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| --- | --- | --- | --- |
| **Solution 1** | | **Solution 2** | |
| Couleur avec bleu de bromothymol | Couleur avec phénolphtaléine | Couleur avec bleu de bromothymol | Couleur avec phénolphtaléine |
|  |  |  |  |
| Caractère acido-basique de la solution 1 | | Caractère acido-basique de la solution 2 | |
|  | |  | |
| Nom de la solution 1 | | Nom de la solution 2 | |
|  | |  | |

* Caractère acido-basique de la solution 3 : ………………………………………………….
* Nom de la solution 3 : …………………………………………………….

**\*Réactions en milieu aqueux :**

* Réaction 1:……………………………………………………………………………
* Réaction 2 :……………………………………………………………………………
* Réaction 3 :……………………………………………………………………………
* Faire un schéma du dispositif de dosage de NaOH par HCl.
* Ecrire les réactions de neutralisation et déduire les couples acido-basiques mis en jeu.

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* La solution contenue dans l’erlenmeyer à la fin du dosage est-elle acide, basique ou neutre ?

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* Calculer la normalité de l’acide acétique CH3COOH (NA).

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* Calculer la normalité de la soude NaOH (NB).

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* Calculer l’incertitude sur la normalité ∆NA.

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* Calculer l’incertitude sur la normalité ∆NB.

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* Déduire la concentration molaire de l’acide acétique et de la soude.

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* Calculer la concentration massique de l’acide acétique et de la soude.

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* Quel est le rôle d’un indicateur coloré dans un dosage ?

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* Le rajout de l’eau modifie-t-il la position de l’équivalence ?

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* **Conclusion**

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